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that the volume of a

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gas varies directly with the Kelvin temperature, assuming that pressure is constant. We use the following formulas: - or $x = x T K oc + 273$ solve the following problems assuming a constant pressure. 2. 3. 4, 5, 6

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changes at constant temperature 1.

Bacteria produce methane gas in sewage-treatment plants. This gas is often captured or burned. If a bacterial culture produces 60.0 mL of methane gas at 700.0 mm Hg, what volume would be produced at 760.0 mm Hg? 2.

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zero, Boyle's law, Charles' law, Kelvin scale, pressure Prior Knowledge Question (Do this BEFORE using the Gizmo.) A small helium tank measures about two feet (60 cm) high. Yet it can fill over 50 balloons! How

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states that the volume
of a gas varies
inversely with its
pressure if
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into one equation. Each of these laws can be derived from this law.
Guy-Lussac's Law $PV/T = k$
 $P_1 V_1 T_2 = P_2 V_2 T_1$
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